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Title : Determination of Silver by Gravimetric Analysis

Objectives

- To determine the amount of silver in grams/ml.
- To calculate the weight percent of silver in the salt and to compare with the theoretical value.
- To study the gravimetric analysis method to determine the assigned ion in an unknown salt.

Introduction and background

Gravimetric analysis is a technique through which the amount of an analyte (the ion being analysed) can be determined through the measurement of mass. Certain ions are precipitated as sparingly soluble compounds containing the analyte. The mass of the precipitate is then determined. A weighed sample is dissolved after which an excess of a precipitating agent is added. The precipitate which forms is filtered, dried and weighed. From the mass and known composition of the precipitate, the amount of the original ion can be determined.

For successful determination in the article must be met. First, the chemical substance must be completely precipitated. In most determinations the precipitate is of such low solubility that losses from dissolution are negligible. An additional factor is the "accuracy of filter", the factor values for solubility of the precipitate. When Ag⁺ is precipitated by addition of Cl⁻

$$Ag^+ + Cl^- \rightarrow AgCl(s)$$

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When the pH is buffered in the range of 5 to 9, the formation of the red chelate occurs quantitatively in a solution.. The Gravimetric Estimation of Nickel: The nickel is precipitated as nickel dimethyl glyoxime by adding alcoholic solution of dimethyl glyoxime $C_4H_6(NO)_2$ and then adding a slight excess of aqueous ammonia solution.

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